Name _____ **Class** _____ **Date** _____

Draw a ring around a number of stars for each statement. If you are very confident about a statement, draw your ring around all the stars. If you do not know anything about a statement do not draw a ring.

Topic	At the end of the unit:	
8Fa		
	Write and identify the chemical symbols for elements.	* * * * *
	Describe Dalton's ideas on atoms, molecules, elements and compounds.	* * * * *
	Write word equations for simple and complex chemical reactions.	* * * * *
	Describe elements using physical properties.	* * * * *
8Fb		
	Explain the difference between physical and chemical changes and properties.	* * * * *
	Use observations to decide whether a chemical reaction has taken place.	* * * * *
	Write and interpret chemical formulae for compounds.	* * * * *
	Explain what happens during chemical reactions using atomic theory.	* * * * *
	Carry out calculations involving the masses of reactants and products.	* * * * *
8Fc		
	Use the periodic table to find symbols and elements with similar properties.	* * * * *
	Identify alkali metals, halogens and noble gases in the periodic table and describe their typical properties.	* * * * *
	Describe how Mendeleev arranged the elements in the periodic table and made predictions about elements.	* * * * *
	Describe how the modern periodic table is arranged.	* * * * *
8Fc Working scientifically		
	Explain what is meant by an anomalous result (outlier).	* * * * *
	Identify anomalous results and the range of readings in data.	* * * * *
	Suggest scientific reasons for anomalous results (outliers).	* * * * *
8Fd		
	State what happens at the melting/freezing/boiling point of a substance.	* * * * *
	Use melting, freezing and boiling points to predict state.	* * * * *
	Identify metals and non-metals by their properties and position in the periodic table.	* * * * *
	Describe and identify trends in physical properties in the periodic table.	* * * * *
8Fe		
	Describe the reactions of metals and non-metals with oxygen and water.	* * * * *
	Compare the properties of metal and non-metal oxides.	* * * * *
	Identify trends in chemical properties within a group.	* * * * *
	Make predictions about chemical properties and reactivity in a group.	* * * * *